

Carolina Biological Supply Company

Keep Calm and Chemistry On: Successful Lab Activities for the New Chemistry Teacher

#carolinaNSTA

Workshop Overview

- **View and learn how to perform engaging demos**
- **Discuss phenomena and students connecting to the lesson**
- **Learn how to repurpose demos as 3D-instruction phenomena**
- **Showcase Carolina's chemistry kit offerings:**
 - **Carolina ChemKits®**
 - **Carolina Chemonstrations®**
 - **Carolina Digital Resources**



Building Toward 3-Dimensional Learning

Scientific and Engineering Practices	Disciplinary Core Ideas	Crosscutting Concepts
<ul style="list-style-type: none">• Developing and using models• Planning and carrying out investigations• Analyzing and interpreting data• Constructing explanations	<p>PS 1: Matter and its interactions</p> <p>PS 3: Energy</p>	<ul style="list-style-type: none">• Cause and effect: Mechanism and explanation• Scale, proportion, and quantity• Systems and system models• Energy and matter: Flows, cycles, and conservation¹

1. NGSS Lead States, *Next Generation Science Standards: For States, By States* (Washington, DC: The National Academies Press, 2013), retrieved from www.nextgenscience.org or ngss.nsta.org.

Workshop Safety

Featured Digital Content

Personal Safety Video



Carolina General Safety Video

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Whoosh Bottle Phenomenon or Phenomenal?



Whoosh bottle video webinar segment 4:41–6:22

Aspects of Phenomena

1. Real world
2. Relevant
3. Potentially puzzling
4. Specific (has context)
5. Meaningful
6. Instructionally productive



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Whoosh Bottle Video Segment 6:18–8:03

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Aspects of Phenomena

1. Real world

Connects to an event or experience

2. Relevant

The demo itself is relevant or can be made relevant to the student

3. Potentially puzzling

Doesn't give itself away; students have something to figure out

4. Specific (has context)

More than student entertainment

5. Meaningful

Students will see value in figuring out the phenomenon

6. Instructionally productive

Builds toward performance expectations

Workshop Reminders

- PowerPoint® presentation for this workshop available at www.carolina.com
- Chemistry webinar available at www.carolina.com; type “webinar” in the search bar



Highly Visual Chemistry Phenomena for 3D Instruction - Web

44:30

Presenters: Matt Bostic and Chris Petersen, Carolina Product Developers

Grades: 6-12

What are the aspects of good phenomena? How can you redesign classic, tried-and-true chemistry demos to support 3-dimensional learning? Find out in this webinar.

- Handout includes all demos and activities from the workshop, including Whoosh Bottle

Demo Golden Beaker

An extremely endothermic reaction

Carolina Chemical Demonstrations in progress

Solid barium hydroxide octahydrate and solid ammonium chloride



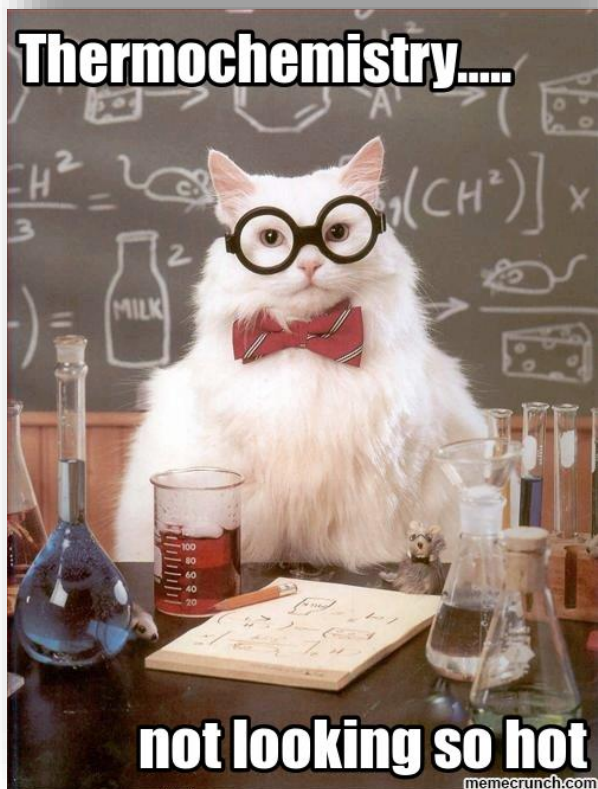
This reaction may reach a temperature as low as -30°C .

Featured Digital Content

- **Student guide copy master**
- **Fill-in answer sheets**
- **Editable assessment questions**
- **Whiteboard resources**

Demo: Frozen Beaker

An extreme endothermic reaction



Curriculum connections:

- Laws of thermodynamics
- Energy changes
- Endothermic vs. exothermic reactions

Possible discussion questions:

- What would be some practical real-world applications of an endothermic reaction?
- Could you measure the heat of reaction in this demo? If not, why?
- Are heat and temperature the same thing?

Lab: Mystery Chemical Reactions

	sodium phosphate	iron(III) chloride	copper(II) sulfate	potassium iodide	lead(II) nitrate	sodium carbonate	silver nitrate	calcium chloride	sodium hydroxide
hydrochloric acid									
sodium hydroxide									
calcium chloride									
silver nitrate									
sodium carbonate									
lead(II) nitrate									
potassium iodide									
copper(II) sulfate									
iron(III) chloride									

Explore scientific phenomena

**Identify visible signs of reaction
(precipitate, gas, and/or color change)**

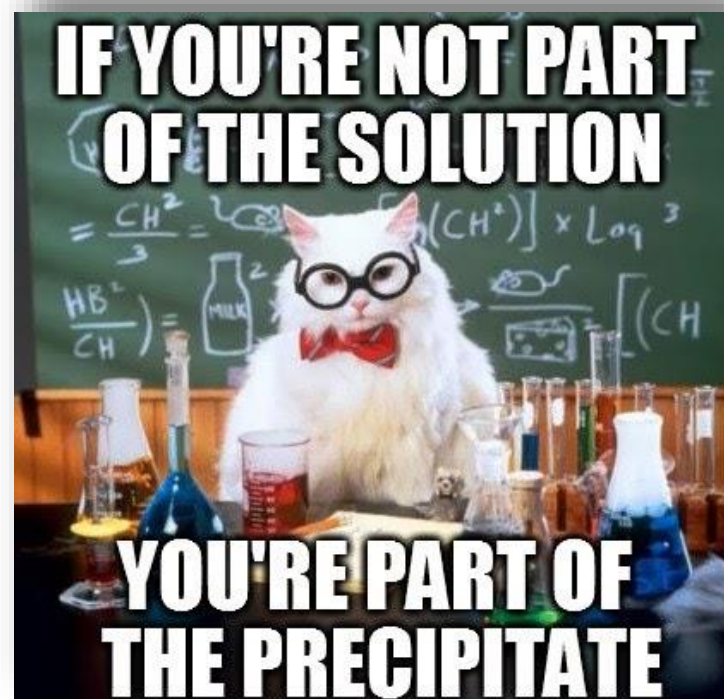
**Microscale chemistry benefits
(save time and money; reduce waste)**

Featured Digital Content

- Interactive lessons

Lab: Mystery Chemical Reactions

	sodium phosphate	iron(III) chloride	copper(II) sulfate	potassium iodide	lead(II) nitrate	sodium carbonate	silver nitrate	calcium chloride	sodium hydroxide
hydrochloric acid		yellow precipitate	light blue precipitate		white precipitate		white precipitate		
sodium hydroxide		yellow precipitate	dark blue precipitate		white precipitate		brown precipitate		
calcium chloride		yellow precipitate	light blue precipitate				white precipitate		
silver nitrate		yellow precipitate	light blue precipitate	yellow precipitate					
sodium carbonate		yellow precipitate	light blue precipitate		white precipitate				
lead(II) nitrate		yellow precipitate	light blue precipitate	yellow precipitate					
potassium iodide		red precipitate	light blue precipitate						
copper(II) sulfate		orange precipitate							
iron(III) chloride		yellow precipitate							



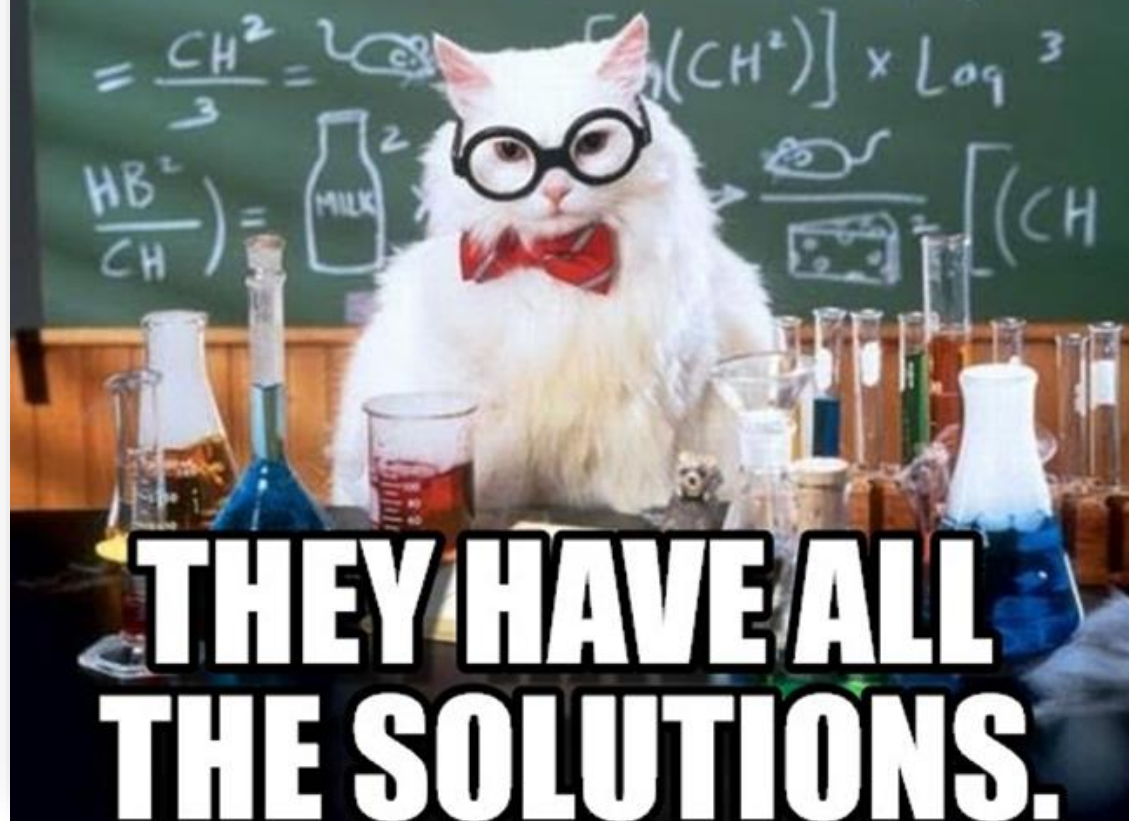
Lab: Balancing Chemical Equations

A tactile introduction to stoichiometry

- Understand the Law of Conservation of Mass
- Understand the difference between coefficients and subscripts in chemical equations



**WHY ARE CHEMISTS GREAT
FOR SOLVING PROBLEMS?**

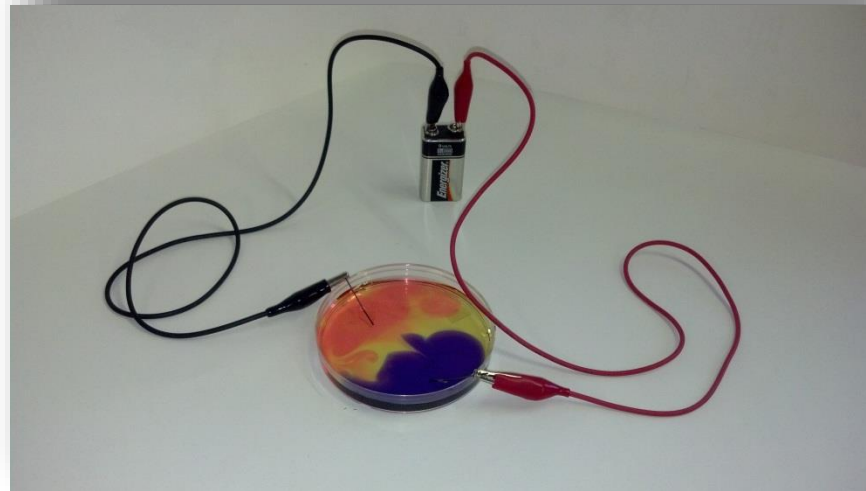


**THEY HAVE ALL
THE SOLUTIONS.**

Lab: Petri Dish Electrolysis

Can we destroy water?

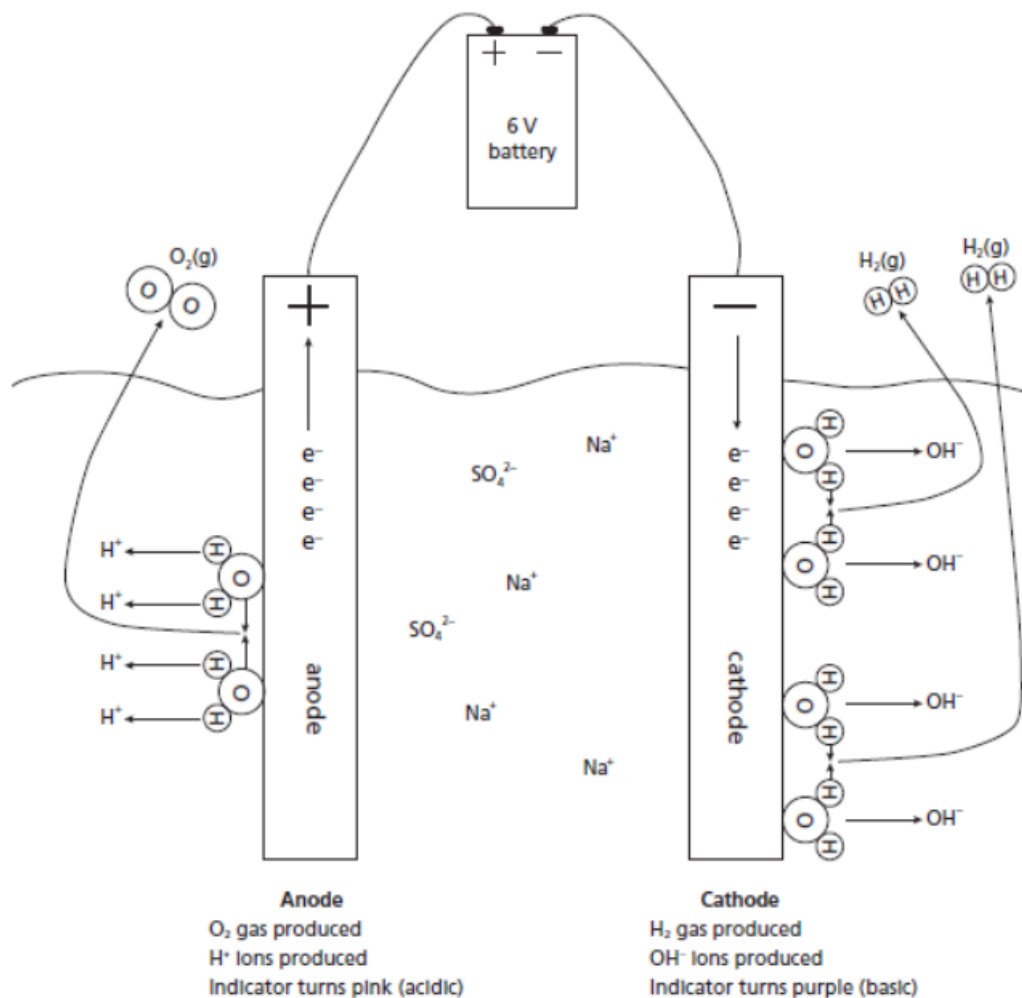
- Explore redox reactions
- Discover electrolysis and identify the products of water electrolysis
- Learn about electrolytes and pH indicators
- Observe visible signs of decomposition



Lab: Petri Dish Electrolysis

Example Particle Level Explanation Diagram

- **Color** = pH at electrode
 - Purple is basic, OH^- ions produced
 - Pink/orange is acidic, H^+ ions produced
- **Bubbles** at electrodes
 - Higher volume is H_2
 - Lower volume is O_2
- **Reduction** happens at the cathode; H^+ is reduced to H_2
- **Oxidation** happens at the anode; OH^- is oxidized to O_2
- **Put together the evidence:**
 - Purple electrode with more bubbles = Cathode
 - Pink/orange electrode with fewer bubbles = Anode



Workshop Kit Review



**Carolina
Chemonstrations®:
Beaker Freezer
Item # 840378**



**Carolina
ChemKits®:
Mystery Chemical
Reactions
Item #840660**

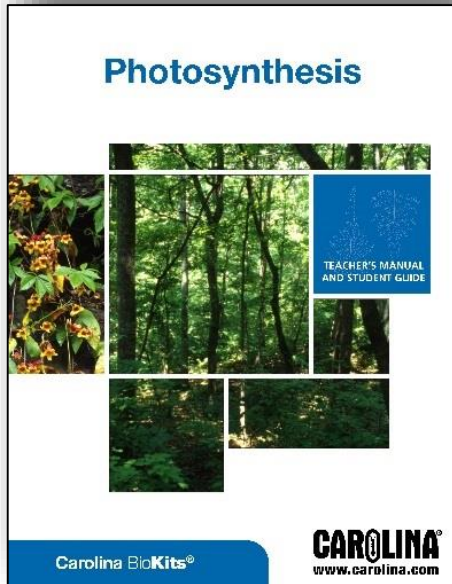


**Carolina
ChemKits®:
Petri Dish
Electrolysis
Item #840830**



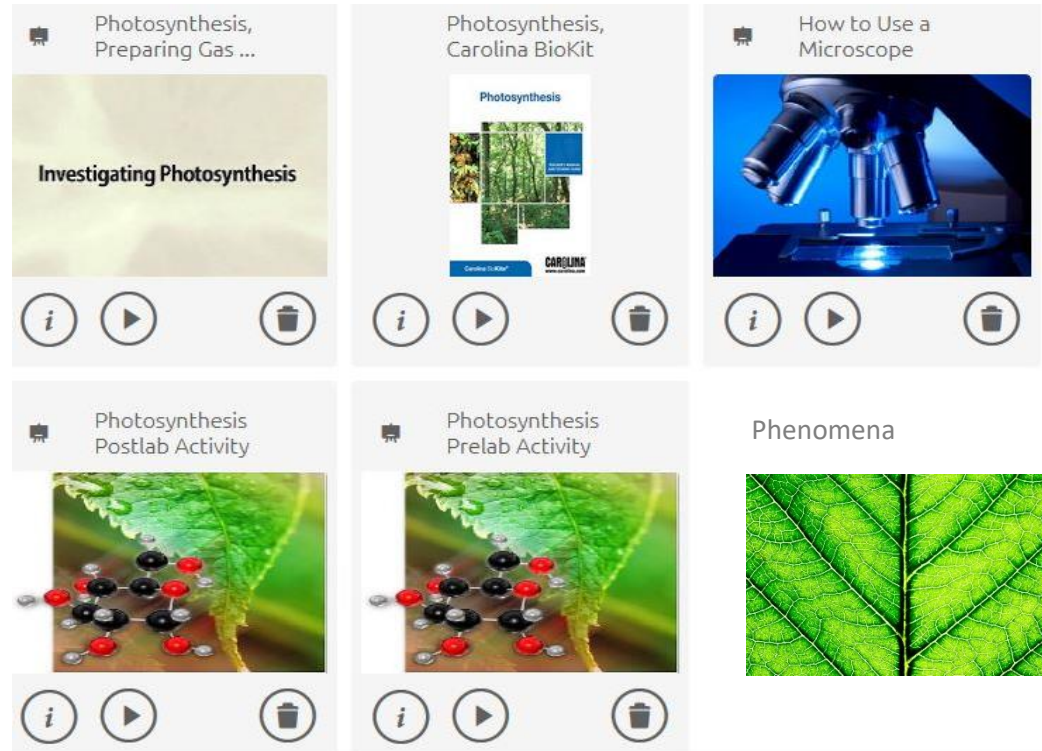
**Carolina
ChemKits®:
Balancing Chemical
Equations
Item #840656**

Kit Manuals & Digital Content Highlights



Digital Assets

- **Start with phenomena**
- **"Stop & Think," scaffolding**
- **Teacher tips**
- **Address NGSS**
- **Easy-to-follow format**



- **Digital teacher's manual**
- **Student guide copy master**
- **Fill-in answer sheets**
- **Editable assessment questions**
- **Whiteboard resources**
- **Phenomena videos**
- **Procedure/intro videos**
- **Pre-lab activity**
- **Post-lab activity**
- **Assessments**
- **Simulations**
- **Safety videos**

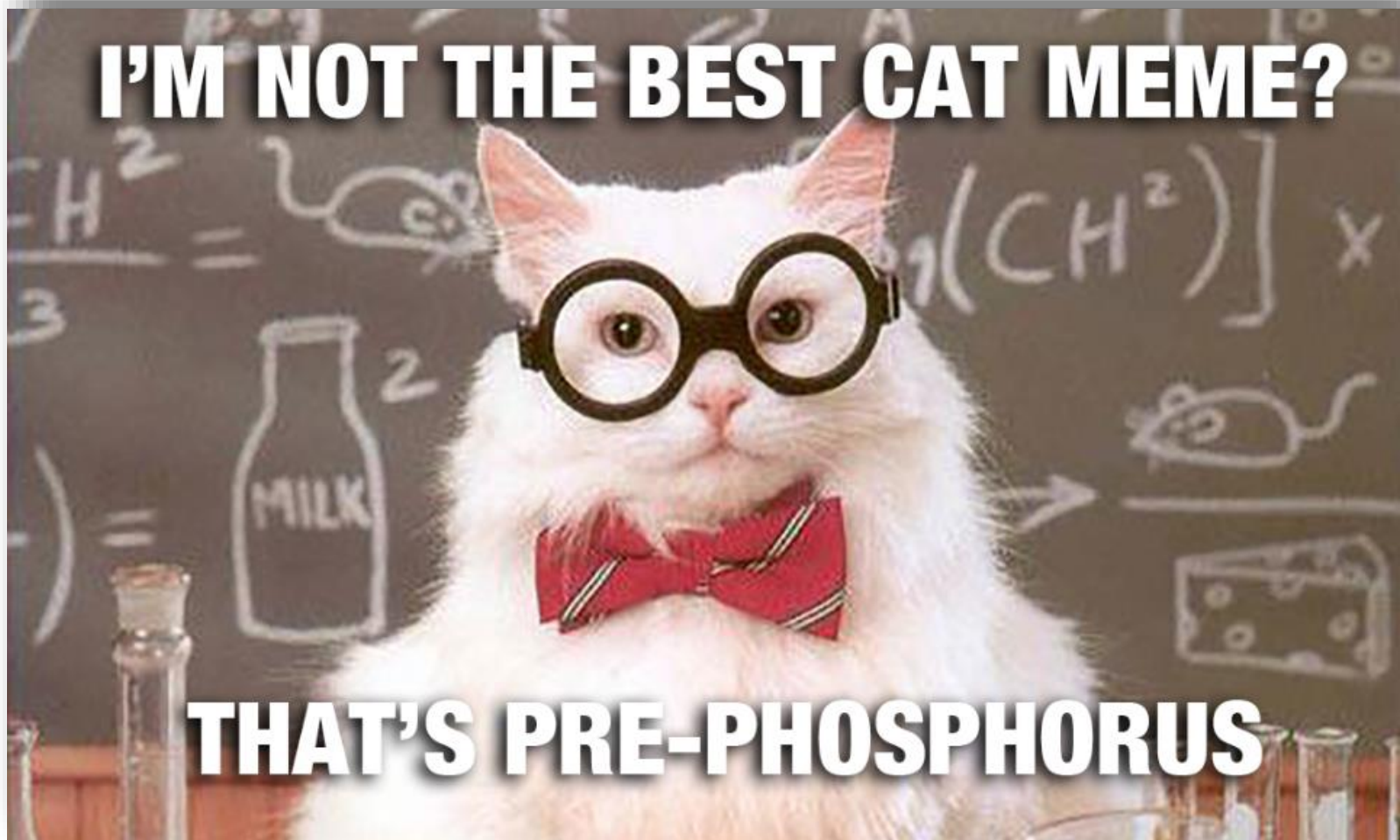
Introducing . . .



Carolina® Spectroscopy Chamber (item #653355)

- Works with Android™ and iOS® devices
- Uses free RGB color analyzer apps
- Does not require special optical glass tubes
- Affordable introduction to spectral analysis

I'M NOT THE BEST CAT MEME?



THAT'S PRE-PHOSPHORUS

Carolina Offers Free Resources to Support Teachers

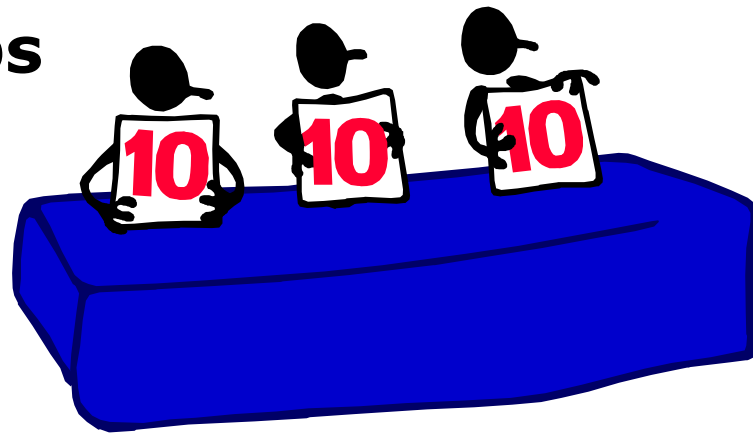


carolina tips®



Evaluations: Share Your Thoughts!

**We are striving to
make our workshops
great!**



**Please evaluate this session and presenter
on a scale from 0 to 10 (10 = best).**

**Please help us reset the room
by gathering your belongings
and exiting between sessions.**

THANK YOU!